

## C20-CHPP -303

## 7277

# BOARD DIPLOMA EXAMINATION, (C-20) OCTOBER/NOVEMBER—2023

#### **DCHE - THIRD SEMESTER EXAMINATION**

ORGANIC AND PHYSICAL CHEMISTRY

Time: 3 hours [ Total Marks: 80

PART—A

 $3 \times 10 = 30$ 

**Instructions:** (1) Answer **all** questions.

- (2) Each question carries three marks.
- (3) Answers should be brief and straight to the point and shall not exceed five simple sentences.
- 1. Write the IUPAC names of the following molecules:

$$\begin{array}{c} \operatorname{CH_3} \\ | \\ \text{(a)} \quad \operatorname{CH_3---} \operatorname{CH} -- \operatorname{CH_3} \\ | \\ \operatorname{CH_3} \end{array}$$

- 2. Explain chlorination of methane.
- **3.** What is HVZ (Hell-Volhard-Zelinsky) reaction?
- **4.** What is Aromaticity? Explain.
- **5.** Write any three differences between metallic conductors and electrolytic conductors.

- **6.** Define the terms (a) Chemical Equivalent and (b) Faraday's constant.
- **7.** Write any three characteristic properties of chemical equilibrium.
- **8.** What are homogeneous and heterogeneous equilibriums? Give examples.
- **9.** State the 1st Law of Thermodynamics. Give mathematical formula for closed system.
- **10.** What is heat of neutralization? Give one example.

#### PART—B

 $10 \times 5 = 50$ 

**Instructions:** (1) Answer any **five** questions.

- (2) Each question carries ten marks.
- (3) Answers should be comprehensive and criterion for valuation is the content but not the length of the answer.
- **11.** (a) Explain about types of stereoisomerism with examples.

#### (OR)

- (b) How do you prepare ethylene from  $C_2H_5OH$  and  $C_2H_5Cl$ ? How does it react with (i) Cold alk.  $KMnO_4$  and (ii)  $Br_2/CCl_4$ ?
- **12.** (a) Describe the manufacture of Ethyl Alcohol from Molasses.

#### (OR)

(b) How do you prepare (i) Chloroform from Ethyl Alcohol and (ii) Acetic acid from Ester?

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**13.** (a) How do you prepare Aniline from Nitro Benzene? Write any three chemical properties of Aniline.

## (OR)

- (b) Write any one method of preparation and two chemical properties of phenol.
- **14.** (a) State and explain Faraday's 2nd law. During the electrolysis of an aqueous solution of NaCl, 10·3 g of Cl<sub>2</sub> is liberated at anode. Calculate the weight of H<sub>2</sub> liberated from acidulated water, if same quantity of electricity is used in the electrolysis of NaCl.

## (OR)

- (b) What are Electrolytes? Explain different types of electrolytes with examples.
- **15.** (a) State and explain law of Mass Action.

# (OR)

- (b) Derive the equilibrium constants for the following reactions:
  - (i)  $N_2$ +3 $H_2 \rightleftharpoons 2NH_3$
  - (ii)  $2SO_2+O_2 \rightleftharpoons 2SO_3$

**PART—C**  $10 \times 1 = 10$ 

**Instructions:** (1) Answer the following question.

- (2) The question carries **ten** marks.
- (3) Answers should be brief and straight to the point and shall not exceed five simple sentences.

## **16.** Explain the following terms :

- (a) Aldol Condensation
- (b) Fehling Test
- (c) Silver Mirror Test

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